Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

9-Aminoacridinium nitrate monohydrate

Mehrdad Pourayoubi,[a]* Hossein Eshtiagh-Hosseini,[a] Somayyeh Sanaei Ataabadi,[b] Teresa Mancilla Percino[c] and Marco A. Leyva Ramírez[c]

[a]Department of Chemistry, Ferdowsi University of Mashhad, Mashhad, 91779, Iran, [b]Department of Chemistry, Islamic Azad University, Shahr-e-Rey Branch, Tehran, Iran, and [c]Departamento de Química, Centro de Investigación y de Estudios Avanzados del Instituto Politécnico Nacional, Apartado Postal 14-740, 07000 México, DF, Mexico
Correspondence e-mail: mehrdad_pourayoubi@yahoo.com

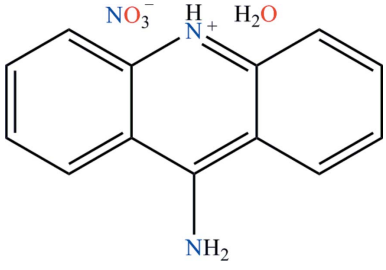
Received 6 January 2011; accepted 31 January 2011

Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.002 Å; R factor = 0.045; wR factor = 0.132; data-to-parameter ratio = 14.0.

The pyridine N atom of the cation in the title hydrated salt, $C_{13}H_{11}N_2^+\cdot NO_3^-\cdot H_2O$, is protonated; the N atom of the NH_2 group shows a planar conformation. The former N atom is hydrogen bonded to a water molecule. The amino group is involved in three N—H···O hydrogen bonds with two neighboring nitrate anions. The water molecule is hydrogen bonded to two adjacent nitrate anions. In the crystal, this results in a layered network.

Related literature

For the structure of 9-aminoacridine hydrochloride monohydrate, see: Talacki *et al.* (1974). For positive-charge-assisted hydrogen bonds, see: Gilli *et al.* (1994).



Experimental

Crystal data

$C_{13}H_{11}N_2^+\cdot NO_3^-\cdot H_2O$
M_r = 275.26
Triclinic, $P\overline{1}$
a = 6.8556 (2) Å
b = 10.0532 (2) Å
c = 10.5912 (3) Å
α = 117.016 (1)°
β = 94.138 (1)°
γ = 97.995 (1)°
V = 636.36 (3) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.11 mm^{-1}
T = 293 K
0.75 × 0.75 × 0.45 mm

Data collection

Nonius KappaCCD diffractometer
Absorption correction: multi-scan (Blessing, 1995)
T_{min} = 0.923, T_{max} = 0.953
8945 measured reflections
2822 independent reflections
2054 reflections with $I > 2\sigma(I)$
R_{int} = 0.028

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.045
$wR(F^2)$ = 0.132
S = 1.04
2822 reflections
201 parameters
5 restraints
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.26 e Å^{-3}
$\Delta\rho_{min}$ = −0.22 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N2—H2A···O1^{i}	0.93 (1)	2.23 (2)	3.0619 (17)	149 (1)
N2—H2A···O3^{i}	0.93 (1)	2.30 (2)	3.0662 (16)	140 (1)
N2—H2B···O2^{ii}	0.90 (1)	2.07 (1)	2.9123 (15)	157 (2)
O4—H4A···O3^{iii}	0.91 (2)	2.03 (2)	2.9147 (18)	164 (2)
N1—H1···O4	0.89 (1)	1.91 (1)	2.7867 (15)	170 (2)
O4—H4B···O1	0.90 (2)	2.01 (2)	2.9058 (18)	173 (2)
O4—H4B···O2	0.90 (2)	2.64 (2)	3.2039 (19)	122 (2)

Symmetry codes: (i) $x, y-1, z-1$; (ii) $-x, -y+1, -z+1$; (iii) $x+1, y, z$.

Data collection: *COLLECT* (Nonius, 2001); cell refinement: *SCALEPACK* (Otwinowski & Minor, 1997); data reduction: *DENZO* (Otwinowski & Minor, 1997) and *SCALEPACK*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *Mercury* (Macrae *et al.*, 2006); software used to prepare material for publication: *WinGX* (Farrugia, 1999) and *enCIFer* (Allen *et al.*, 2004).

Support of this investigation by Ferdowsi University of Mashhad and the Islamic Azad University Shahr-e-Rey Branch is gratefully acknowledged.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG5102).

References

Allen, F. H., Johnson, O., Shields, G. P., Smith, B. R. & Towler, M. (2004). *J. Appl. Cryst.* **37**, 335–338.
Blessing, R. H. (1995). *Acta Cryst.*A**51**, 33–38.
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
Gilli, P., Bertolasi, V., Ferretti, V. & Gilli, G. (1994). *J. Am. Chem. Soc.* **116**, 909–915.
Macrae, C. F., Edgington, P. R., McCabe, P., Pidcock, E., Shields, G. P., Taylor, R., Towler, M. & van de Streek, J. (2006). *J. Appl. Cryst.* **39**, 453–457.
Nonius (2001). *COLLECT*. Nonius BV, Delft, The Netherlands.
Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Talacki, R., Carrell, H. L. & Glusker, J. P. (1974). *Acta Cryst.* B**30**, 1044–1047.

supplementary materials

Acta Cryst. (2011). E**67**, o565 [**doi:10.1107/S1600536811003953**]

9-Aminoacridinium nitrate monohydrate

M. Pourayoubi, H. Eshtiagh-Hosseini, S. Sanaei Ataabadi, T. Mancilla Percino and M. A. Leyva Ramírez

Comment

In a previous work, the crystal structure of 9-aminoacridine hydrochloride monohydrate (Talacki *et al.*, 1974) has been investigated. Here, we report on the crystal structure of title hydrated salt, $C_{13}H_{11}N_2^+.NO_3^-.H_2O$ (Fig. 1).

In 9-amino-acridinium cation, the heteroatom N1 and the nitrogen atom of NH_2 unit (N2) have a sp^2 character. The C1—N1—C13 angle is 122.68 (11)°; the fused tricyclic system is essentially planar.

The protonated pyridine nitrogen atom is involving in a positive charge assisted (Gilli *et al.*, 1994) N—H···O hydrogen bond with a neighboring H_2O molecule (N1···O4 = 2.7867 (15) Å). Moreover, the water molecule forms two O—H···O hydrogen bonds (O···O = 2.9058 (18) & 2.9147 (18) Å) with two adjacent NO_3^- anions; also, the weak hydrogen bond O4—H4B···O2 (O4···O2 = 3.2039 (19) Å) may be considered which has not influence on the pattern of crystal packing. The NH_2 unit of cation cooperates in three N—H···O hydrogen bonds (N···O = 2.9123 (15), 3.0619 (17) and 3.0662 (16) Å), with two neighboring nitrate anions. Cations, anions and water molecules are hydrogen bonded in a 2-D arrangement (Fig. 2).

Experimental

The title hydrated salt was obtained fortuitously from the reaction between 9-aminoacridine and $Fe(NO_3)_3.9H_2O$ in CH_3OH as follows: To a solution of 9-aminoacridine (0.194 g, 1 mmol) in CH_3OH (5 ml), a solution of $Fe(NO_3)_3.9H_2O$ (0.202 g, 0.5 mmol) in CH_3OH (5 ml) was added at 343 K. After 1 h stirring, the solid was filtered; the crystals were obtained from methanolic solution after a slow evaporation at room temperature.

Refinement

The hydrogen atom of NH group and those of water molecule were found in difference Fourier synthesis.The NH H atoms were restrained to 0.90 A and the refinement give good values. The H atoms in the water molecule were refined with a restraint of 1.00 A for a ideal distance OH and obtained acceptable values. The H(C) atom positions were calculated. All hydrogen atoms were refined in isotropic approximation in riding model with the U_{iso}(H) parameters equal to 1.2 U_{eq}(Ci), for methyl groups equal to 1.5 U_{eq}(Cii), where U(Ci) and U(Cii) are respectively the equivalent thermal parameters of the carbon atoms to which corresponding H atoms are bonded.

Figures

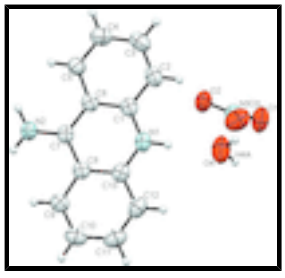


Fig. 1. Molecular view with the atom labeling scheme, displacement ellipsoids are drawn at the 50% probability level. H atoms are represented as small spheres of arbitrary radii.

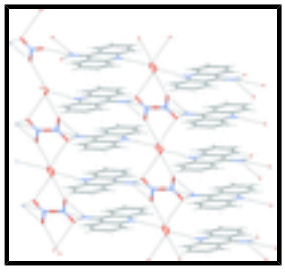


Fig. 2. Partial packing of cations, anions and water molecules in the title hydrated salt. H bonds are shown as dashed lines.

9-Aminoacridinium nitrate monohydrate

Crystal data

$C_{13}H_{11}N_2^+ \cdot NO_3^- \cdot H_2O$	$Z = 2$
$M_r = 275.26$	$F(000) = 288$
Triclinic, $P\overline{1}$	$D_x = 1.437$ Mg m^{-3}
Hall symbol: -P 1	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
$a = 6.8556$ (2) Å	Cell parameters from 600 reflections
$b = 10.0532$ (2) Å	$\theta = 1$–$14°$
$c = 10.5912$ (3) Å	$\mu = 0.11$ mm^{-1}
$\alpha = 117.016$ (1)°	$T = 293$ K
$\beta = 94.138$ (1)°	Block, colourless
$\gamma = 97.995$ (1)°	0.75 × 0.75 × 0.45 mm
$V = 636.36$ (3) Å^3	

Data collection

Nonius KappaCCD diffractometer	2822 independent reflections
Radiation source: fine-focus sealed tube	2054 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.028$
CCD rotation images, thick slices scans	$\theta_{max} = 27.6°$, $\theta_{min} = 3.0°$
Absorption correction: multi-scan (Blessing, 1995)	$h = -8 \rightarrow 8$
$T_{min} = 0.923$, $T_{max} = 0.953$	$k = -12 \rightarrow 13$
8945 measured reflections	$l = -13 \rightarrow 13$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.045$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.132$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.04$	$w = 1/[\sigma^2(F_o^2) + (0.0799P)^2 + 0.0299P]$ where $P = (F_o^2 + 2F_c^2)/3$
2822 reflections	$(\Delta/\sigma)_{max} < 0.001$
201 parameters	$\Delta\rho_{max} = 0.26$ e Å^{-3}
5 restraints	$\Delta\rho_{min} = -0.22$ e Å^{-3}

Special details

Geometry. All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
C1	0.28645 (17)	0.69390 (14)	0.48088 (14)	0.0392 (3)
C2	0.3072 (2)	0.82982 (15)	0.46976 (16)	0.0491 (3)
H2	0.3418	0.9237	0.5518	0.059*
C3	0.2764 (2)	0.82263 (16)	0.33871 (17)	0.0548 (4)
H3	0.2900	0.9123	0.3317	0.066*
C4	0.2245 (2)	0.68227 (17)	0.21357 (17)	0.0545 (4)
H4	0.2054	0.6794	0.1245	0.065*
C5	0.20191 (19)	0.54989 (15)	0.22257 (14)	0.0456 (3)
H5	0.1663	0.4572	0.1393	0.055*
C6	0.23215 (17)	0.55232 (13)	0.35736 (13)	0.0377 (3)
C7	0.20689 (17)	0.41608 (13)	0.37264 (13)	0.0368 (3)
C8	0.23830 (16)	0.43023 (13)	0.51410 (13)	0.0369 (3)
C9	0.21635 (19)	0.30332 (15)	0.54077 (15)	0.0438 (3)
H9	0.1779	0.2057	0.4642	0.053*
C10	0.2506 (2)	0.32195 (17)	0.67683 (16)	0.0513 (4)
H10	0.2353	0.2375	0.6924	0.062*
C11	0.3088 (2)	0.46790 (18)	0.79260 (16)	0.0554 (4)

H11	0.3326	0.4797	0.8850	0.066*
C12	0.3312 (2)	0.59339 (17)	0.77252 (14)	0.0512 (4)
H12	0.3695	0.6900	0.8507	0.061*
C13	0.29601 (17)	0.57631 (14)	0.63302 (13)	0.0393 (3)
N1	0.31813 (16)	0.70224 (12)	0.61313 (12)	0.0437 (3)
N2	0.15562 (19)	0.28113 (13)	0.25875 (12)	0.0507 (3)
N3	0.0612 (2)	1.10579 (12)	0.88933 (12)	0.0521 (3)
O1	0.22383 (17)	1.19337 (12)	0.95048 (12)	0.0721 (4)
O2	0.04483 (18)	1.00530 (11)	0.76309 (11)	0.0674 (3)
O3	−0.08334 (18)	1.12104 (15)	0.95548 (12)	0.0739 (4)
O4	0.4979 (2)	0.98619 (13)	0.83821 (14)	0.0775 (4)
H1	0.361 (2)	0.7928 (17)	0.6889 (16)	0.062 (5)*
H2A	0.133 (2)	0.2695 (19)	0.1668 (16)	0.061 (4)*
H2B	0.126 (2)	0.1971 (17)	0.2686 (18)	0.067 (5)*
H4A	0.626 (2)	1.017 (2)	0.883 (2)	0.094 (7)*
H4B	0.422 (3)	1.057 (2)	0.875 (2)	0.098 (7)*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0320 (6)	0.0385 (6)	0.0459 (7)	0.0080 (5)	0.0077 (5)	0.0183 (6)
C2	0.0462 (7)	0.0360 (6)	0.0613 (9)	0.0068 (5)	0.0078 (6)	0.0200 (6)
C3	0.0529 (8)	0.0489 (8)	0.0750 (10)	0.0128 (6)	0.0108 (7)	0.0386 (8)
C4	0.0580 (8)	0.0612 (9)	0.0577 (9)	0.0190 (7)	0.0121 (6)	0.0369 (8)
C5	0.0480 (7)	0.0453 (7)	0.0441 (7)	0.0138 (6)	0.0073 (5)	0.0202 (6)
C6	0.0321 (6)	0.0383 (6)	0.0432 (7)	0.0097 (5)	0.0083 (5)	0.0183 (6)
C7	0.0315 (6)	0.0356 (6)	0.0398 (7)	0.0079 (5)	0.0061 (5)	0.0142 (5)
C8	0.0288 (6)	0.0407 (7)	0.0418 (7)	0.0092 (5)	0.0076 (5)	0.0189 (6)
C9	0.0407 (7)	0.0435 (7)	0.0495 (7)	0.0097 (5)	0.0093 (5)	0.0229 (6)
C10	0.0475 (7)	0.0610 (9)	0.0605 (9)	0.0167 (6)	0.0154 (6)	0.0387 (8)
C11	0.0543 (8)	0.0750 (10)	0.0448 (8)	0.0185 (7)	0.0128 (6)	0.0325 (8)
C12	0.0494 (8)	0.0565 (8)	0.0390 (7)	0.0103 (6)	0.0078 (6)	0.0148 (6)
C13	0.0326 (6)	0.0428 (7)	0.0399 (7)	0.0084 (5)	0.0077 (5)	0.0166 (6)
N1	0.0450 (6)	0.0353 (6)	0.0412 (6)	0.0056 (5)	0.0052 (5)	0.0107 (5)
N2	0.0694 (8)	0.0356 (6)	0.0397 (6)	0.0073 (5)	0.0013 (5)	0.0133 (5)
N3	0.0707 (8)	0.0388 (6)	0.0428 (6)	0.0083 (6)	−0.0041 (6)	0.0182 (5)
O1	0.0757 (8)	0.0523 (6)	0.0587 (7)	−0.0043 (6)	−0.0034 (6)	0.0076 (5)
O2	0.0965 (8)	0.0445 (6)	0.0431 (6)	0.0017 (5)	−0.0006 (5)	0.0100 (5)
O3	0.0681 (7)	0.0950 (9)	0.0604 (7)	0.0244 (6)	0.0096 (6)	0.0358 (7)
O4	0.0701 (8)	0.0541 (7)	0.0758 (8)	0.0067 (6)	−0.0002 (6)	0.0060 (6)

Geometric parameters (Å, °)

C1—N1	1.3641 (18)	C9—H9	0.9300
C1—C6	1.4024 (18)	C10—C11	1.396 (2)
C1—C2	1.4127 (18)	C10—H10	0.9300
C2—C3	1.356 (2)	C11—C12	1.362 (2)
C2—H2	0.9300	C11—H11	0.9300
C3—C4	1.403 (2)	C12—C13	1.4061 (19)

C3—H3	0.9300	C12—H12	0.9300
C4—C5	1.3654 (19)	C13—N1	1.3650 (17)
C4—H4	0.9300	N1—H1	0.887 (14)
C5—C6	1.4163 (18)	N2—H2A	0.925 (14)
C5—H5	0.9300	N2—H2B	0.895 (14)
C6—C7	1.4393 (17)	N3—O2	1.2411 (15)
C7—N2	1.3186 (16)	N3—O1	1.2417 (16)
C7—C8	1.4361 (17)	N3—O3	1.2417 (17)
C8—C13	1.4091 (18)	O4—H4A	0.909 (16)
C8—C9	1.4178 (17)	O4—H4B	0.901 (16)
C9—C10	1.364 (2)		
N1—C1—C6	120.53 (11)	C10—C9—H9	119.4
N1—C1—C2	119.19 (12)	C8—C9—H9	119.4
C6—C1—C2	120.28 (12)	C9—C10—C11	119.94 (13)
C3—C2—C1	119.60 (13)	C9—C10—H10	120.0
C3—C2—H2	120.2	C11—C10—H10	120.0
C1—C2—H2	120.2	C12—C11—C10	121.11 (13)
C2—C3—C4	121.12 (13)	C12—C11—H11	119.4
C2—C3—H3	119.4	C10—C11—H11	119.4
C4—C3—H3	119.4	C11—C12—C13	119.71 (13)
C5—C4—C3	120.01 (13)	C11—C12—H12	120.1
C5—C4—H4	120.0	C13—C12—H12	120.1
C3—C4—H4	120.0	N1—C13—C12	119.63 (12)
C4—C5—C6	120.69 (13)	N1—C13—C8	120.00 (11)
C4—C5—H5	119.7	C12—C13—C8	120.37 (12)
C6—C5—H5	119.7	C1—N1—C13	122.68 (11)
C1—C6—C5	118.30 (11)	C1—N1—H1	118.7 (11)
C1—C6—C7	118.91 (11)	C13—N1—H1	118.6 (11)
C5—C6—C7	122.78 (11)	C7—N2—H2A	122.2 (10)
N2—C7—C8	120.83 (11)	C7—N2—H2B	120.4 (11)
N2—C7—C6	120.48 (11)	H2A—N2—H2B	116.9 (16)
C8—C7—C6	118.70 (11)	O2—N3—O1	119.79 (14)
C13—C8—C9	117.72 (11)	O2—N3—O3	120.94 (13)
C13—C8—C7	119.16 (11)	O1—N3—O3	119.27 (12)
C9—C8—C7	123.11 (11)	H4A—O4—H4B	114 (2)
C10—C9—C8	121.15 (13)		
N1—C1—C2—C3	179.86 (12)	N2—C7—C8—C9	−0.21 (19)
C6—C1—C2—C3	−0.71 (19)	C6—C7—C8—C9	179.70 (10)
C1—C2—C3—C4	−0.1 (2)	C13—C8—C9—C10	−0.24 (18)
C2—C3—C4—C5	0.7 (2)	C7—C8—C9—C10	179.00 (11)
C3—C4—C5—C6	−0.6 (2)	C8—C9—C10—C11	−0.1 (2)
N1—C1—C6—C5	−179.71 (11)	C9—C10—C11—C12	0.4 (2)
C2—C1—C6—C5	0.87 (17)	C10—C11—C12—C13	−0.2 (2)
N1—C1—C6—C7	1.12 (17)	C11—C12—C13—N1	179.87 (12)
C2—C1—C6—C7	−178.31 (10)	C11—C12—C13—C8	−0.1 (2)
C4—C5—C6—C1	−0.23 (19)	C9—C8—C13—N1	−179.64 (10)
C4—C5—C6—C7	178.91 (11)	C7—C8—C13—N1	1.09 (17)
C1—C6—C7—N2	179.90 (11)	C9—C8—C13—C12	0.38 (17)

C5—C6—C7—N2	0.77 (19)	C7—C8—C13—C12	−178.89 (11)
C1—C6—C7—C8	−0.01 (16)	C6—C1—N1—C13	−1.16 (18)
C5—C6—C7—C8	−179.14 (11)	C2—C1—N1—C13	178.27 (10)
N2—C7—C8—C13	179.02 (11)	C12—C13—N1—C1	−179.99 (11)
C6—C7—C8—C13	−1.08 (16)	C8—C13—N1—C1	0.03 (18)

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N2—H2A···O1^{i}	0.93 (1)	2.23 (2)	3.0619 (17)	149.(1)
N2—H2A···O3^{i}	0.93 (1)	2.30 (2)	3.0662 (16)	140.(1)
N2—H2B···O2^{ii}	0.90 (1)	2.07 (1)	2.9123 (15)	157.(2)
O4—H4A···O3^{iii}	0.91 (2)	2.03 (2)	2.9147 (18)	164.(2)
N1—H1···O4	0.89 (1)	1.91 (1)	2.7867 (15)	170.(2)
O4—H4B···O1	0.90 (2)	2.01 (2)	2.9058 (18)	173 (2)
O4—H4B···O2	0.90 (2)	2.64 (2)	3.2039 (19)	122.(2)

Symmetry codes: (i) *x*, *y*−1, *z*−1; (ii) −*x*, −*y*+1, −*z*+1; (iii) *x*+1, *y*, *z*.

Fig. 1

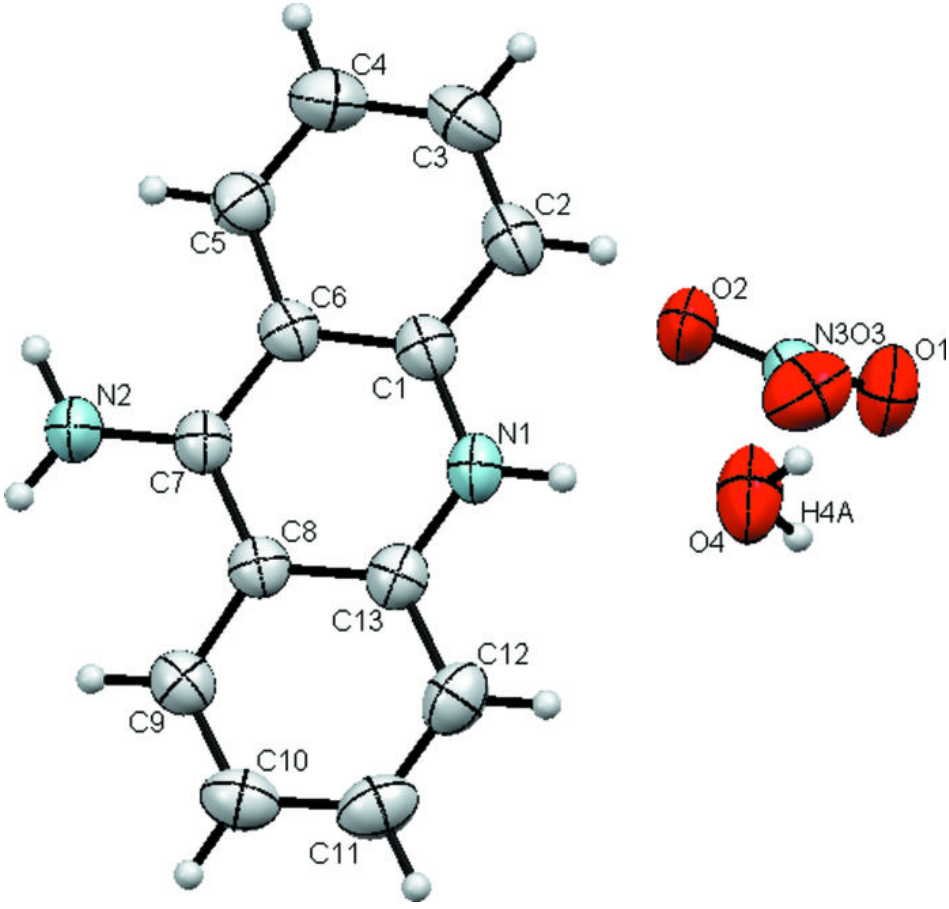


Fig. 2